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| **CHEM-371 ENSTRUMENTAL ANALYSIS LABORATORY****Neutralization Titrations by Conductivity Measurement****LAB REPORT** |
| **Instructors: Assistant:** |
| **NAME & SURNAME: DATE:** **ID: Section ( )** |
| **THE PURPOSE OF THE EXPERIMENT:****PROCEDURE:****DATA AND CALCULATIONS:**1. Titration of strong acid (20 ml and 0.018 M HCl) by strong base (0.2 M NaOH).

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| Added base ( mL) |  Measured conductivity (µs) |
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1. Titration of weak acid ( 20 mL and 0,016 M HAc, Ka = 1x10-5) by strong base ( 0.2 M NaOH)

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| Added base ( mL) |  Measured conductivity (µs) |
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1. **Titration of unknown acid mixture (0,018 M HCl, 0,016 M HAc and mixture total volume = 33 mL) by strong base (0,2 M NaOH)**

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| Added base ( mL) |  Measured conductivity (µs) |
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1. **Titration curve by the 1th titration ( VNaOH versus conductivity)**

milimetrik kağıt ile ilgili görsel sonucu1. **Titration curve by the 2nd titration (VNaOH versus conductivity)**

milimetrik kağıt ile ilgili görsel sonucu1. **Titration curve by the 3rd titration (VNaOH versus conductivity)**

milimetrik kağıt ile ilgili görsel sonucu**CALCULATIONS:****RESULTS AND DISCUSSION**1. **For the 1st Experiment(**Titration of strong acid by strong base**)**

Equivalence point (theoretical): …………….End point (Experimental): …………………..% Error: …………………………….1. **For the 2nd Experiment(**Titration of weak acid by strong base**)**

Equivalence point (theoretical): …………….End point (Experimental): …………………..% Error: …………………………….**CONCLUSION:** |